**Chemistry confidential**

**Instruction to schools**

**Apart from the common laboratory and chemicals, each candidate requires the following;**

**A**

* Solution D -100cm3
* Solution K -100cm3
* Solution N -100cm3
* Solution R -100cm3
* One burette 0-50ml
* One pipette 25ml and pipette filler
* Six dry test tube and test tube holder
* One thermometer
* One metallic spatula
* One boiling tube
* One 10ml measuring cylinder
* 250ml plastic beaker
* Distilled water
* 2-conical flask
* About 2g of solid NaHCO3

**B: ACCESS TO**

* Source of heat
* Distilled water.
* Litmus paper
* 2M NaoH
* 2M Ba(No3)2
* 2M NH3(aq)
* 2M NaCl
* Phenolphthalein
* Acidified potassium dichromate
* Bromine water

**C) FURTHER INFORMATION**

Solid J - Maleic acid (0.5g)

Solid D hydrated Zncl2 (0.5g)

Solution D -Sulphuric (VI) acid 1M

Solution R (NaoH) 1M

Solution N prepared by dissolving 100cm3 concentration of HCL in 400cm3 concentration of HCl in 400cm3 distilled water then top up to make one liter (100cm3)

Solution K (NaOH) dissolves 40g in 400cm3 made upto litre (100cm3)